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## The use of $^{152}\text{Eu}(\text{III})$ towards the elucidation of $\text{Pu}(\text{III})$ behavior with phosphates

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Phosphate species from natural (e.g. phosphate mineral like apatite) as well anthropogenic (e.g. fertilizers) sources are ubiquitous in the environment [1]. Phosphoric acid is also present in the PUREX process as ultimate degradation product of tributylphosphate [2]. Phosphate species are expected to strongly interact with plutonium ions, plutonium being one of the main environmental pollutants associated with nuclear power generation and nuclear weapon testing [3]. In order to improve the models of Pu chemical behavior in the presence of phosphate species, it is necessary to collect fundamental data on interaction of  $\text{Pu}^{3+}$ ,  $\text{Pu}^{4+}$ ,  $\text{PuO}_2^+$ , and  $\text{PuO}_2^{2+}$  ions with these inorganic ligands. However, due to difficulties in handling Pu (redox properties, radiotoxicity), the aim is first to develop an experimental protocol using analogs in the same oxidation state [4]. The present work focuses on the determination of, on one hand, the speciation of Eu(III) at ultra-trace scale using  $^{152}\text{Eu}$  in the presence of  $\text{H}_2\text{PO}_4^-$  and, on the other hand, of the solubility product of Nd-based rhabdophane.

The partition of  $^{152}\text{Eu}$  in the system HDEHP/Toluene/NaCl/HCl/NaH $_2$ PO $_4$  has been studied at constant ionic strength and temperature. The mean composition of the formed complexes and their associated stability constants have been deduced from the variations of the distribution ratio of  $^{152}\text{Eu}$  as function of phosphate concentration.

Concerning the solubility product of  $\text{PuPO}_4$ , stable rhabdophane  $\text{Nd}_{0.99}\text{Eu}_{0.01}\text{PO}_4 \cdot n\text{H}_2\text{O}$  and the  $^{152}\text{Eu}$  doped one have been synthesized. The stable solid has been characterized by XRD, TGA and SEM, and its dissolution has been proved to be congruent. Dissolution experiments have been then conducted on the doped phosphate compounds reducing the amount of solid (from 50 to 2 mg). The process was monitored by  $\gamma$ -spectrometry. The same behavior was observed in all the studied systems: the equilibrium is reached in about ten days and the solubility product obtained in these conditions agrees with literature data [5]. This validates the use of a small amount of  $\text{PuPO}_4$  in future dissolution experiments.

### References

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