

# Radiation-induced modifications on physico-chemical properties of diluted nitric acid solutions within advanced Spent Nuclear Fuel reprocessing

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Within advanced Spent Nuclear Fuel reprocessing, the effects of ionizing radiation on some physico-chemical properties of 0.25-0.5M HNO<sub>3</sub> solutions have been studied, because of the possible influence on both the separation performances and the fluid-dynamics of the extracting system. Irradiations were performed in air up to 100kGy by two <sup>60</sup>Co-sources, with 0.3kGy/h and 2.5kGy/h dose rates, respectively. Density, viscosity, acidity and nitrate ion concentration were measured before and after irradiation. No modifications of these properties in the dose range considered were observed, therefore no significant effects on the fluid-dynamics of the extracting systems are expected.

**KEYWORDS:** partitioning,  $\gamma$ -radiolysis, nitric acid, density, viscosity, acidity.

## 1. INTRODUCTION

In the last decades, the increasing radioactive waste amount coming from nuclear energy power plants has become such a big deal so as to encourage the development of innovative treatments with the aim to reduce the radiotoxicity and the heat production of the final waste which has to be stored in dedicated repository [1]. One of the possible approaches to be pursued is hydro-metallurgical Partitioning of minor actinides coupled with Transmutation in proper burner reactors in order to obtain shorter-lived or even stable nuclides [2,3], reducing the required storage time from 100000 years to few hundreds of years [4]. The most promising Partitioning strategies are based on a three consecutive steps approach: i) separation of uranium and plutonium from spent nuclear fuel; ii) co-extraction of trivalent actinides and lanthanides; iii) separation of trivalent actinides from trivalent lanthanides [2]. The last step is of capital importance because several lanthanides are distinguished by large neutron capture cross sections [5] and their presence would impact on transmutation efficiency [6]. Since the last step is also the most difficult to be pursued due to the similar chemical behaviour of 4f and 5f elements, several hydrometallurgical processes have been proposed [7,8]. Among them, i-SANEX (innovative Selective ActiNides EXtraction) process is one of the most promising. GANEX (Grouped ActiNides EXtraction) process is a valid alternative approach, since it would allow to separate all valence minor actinides [7].

Since these systems will be employed in highly radioactive environments, mainly due to alpha-emitters present in the process streams, the radiation chemistry of both ligands and diluents will play an essential role in determining extraction efficiency, separation factors and solvent-recycling. Radiation-induced effects would likely include ligand concentration decrease and production of ligand and diluent degradation products [1]. These species would undesirably affect also physico-chemical properties of the solvent, hence altering the process performance due to precipitates, third phases and alterations of density and viscosity [9,10]. It is known from the literature that most of the radiation damage falls on the diluent and, indirectly, on the ligands [11]. Therefore, it is paramount to study the effects of radiolysis on the diluents [12], in order to foresee the possible interactions among the extractant and the diluent by-products as well as to evaluate changes of both the separation performances and the fluid-dynamics of the extracting system. The effects of radiation on concentrated nitric acid have been widely studied in the past because of its employment in the earlier separative processes (e.g. PUREX) [12,13,14], but a lack of knowledge regarding the diluents involved in advanced spent nuclear fuel reprocessing (e.g. diluted nitric acid solutions) remains.

The purpose of this experimental work is to <sup>study in depth</sup> go deeper into the effects of radiation on physico-chemical properties of diluted nitric acid in order to assess whether any modification of important physico-chemical properties occurs, resulting in possible alterations of the fluid-dynamics of the extracting system. Such information would be of capital interest for the application within the newest and promising processes based on hydrophilic extractants. For this purpose density, viscosity, acidity and nitrate anion concentration [NO<sub>3</sub><sup>-</sup>] of diluted nitric acid solutions were systematically measured before and after  $\gamma$ -irradiation, at different dose rates and total absorbed doses.

## 2. EXPERIMENTAL

### 2.1. REAGENTS AND MATERIALS

All the analysed solutions were prepared with concentrated nitric acid (65%, Fluka) and distilled water purified with Millipore ultrapurification system. All reagents used in the analyses are analytical grade and were used without further purifications.

### 2.2. EXPERIMENTAL CONDITIONS

0.25M and 0.5M HNO<sub>3</sub> solutions were prepared from concentrated HNO<sub>3</sub> and stored in the dark in glass vials sealed with plastic lid. Irradiation campaigns were performed in air using two <sup>60</sup>Co-sources with approximately 0.3kGy/h and 2.5kGy/h dose rates respectively, up to total absorbed dose of 50kGy for the lower dose rate irradiation and 100kGy for the higher dose rate irradiation. These irradiation conditions are coherent with the conditions that a SANEX stream should undergo [15]. Following irradiation no volume decrease was measured, hence solvent evaporation is expected to be negligible. Then the samples were kept sealed at 4°C in the dark until further analyses.

## 3. RESULTS AND DISCUSSION

### 3.1. DENSITY

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Density measurements on approximately 2.7mL of fresh and irradiated solutions were performed by *DMA 35N Anton Paar portable density meter*, which is affected by a measuring uncertainty of  $\pm 0.001 \text{ g/cm}^3$ , by keeping the samples in a thermostatic bath with temperature resolution of  $\pm 0.1^\circ\text{C}$ . The samples were analysed at temperatures of interest for the separative process, ranging from 20°C to 45°C, with incremental temperature of 5°C. Each density value is the mean of 10 measurements. The estimation of the measuring uncertainty has been carried out according to Minimum Mean Square Error method, considering the variances of the acquired data and the uncertainty associated to the instrument itself [16].

The results obtained for 0.25M HNO<sub>3</sub> irradiated samples are reported in Fig 1 and in Fig 2. It could be observed that no density radiation-induced variation is noticeable for the absorbed dose range considered at both dose rates within the limit of experimental error. Similar trends were found for irradiated 0.5M HNO<sub>3</sub> solutions.

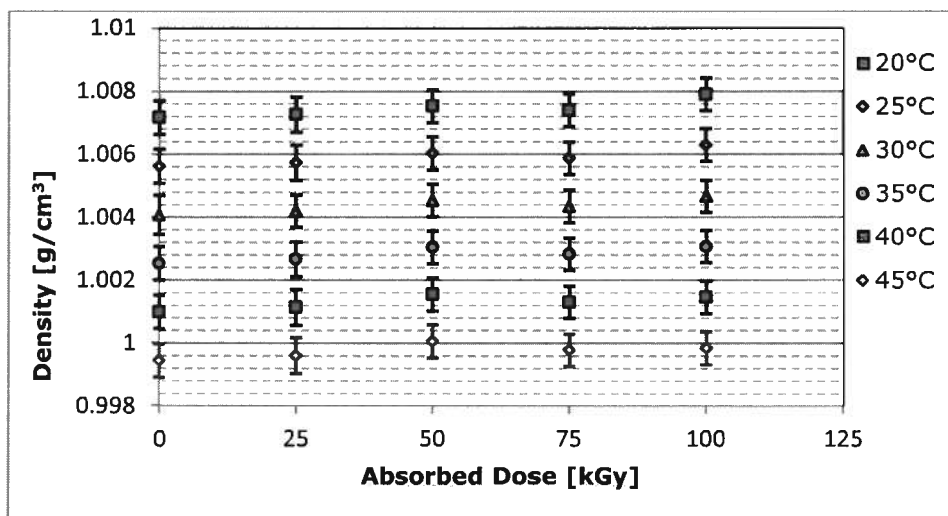


Fig 1 Density of 0.25M HNO<sub>3</sub> as a function of absorbed dose at high dose rate

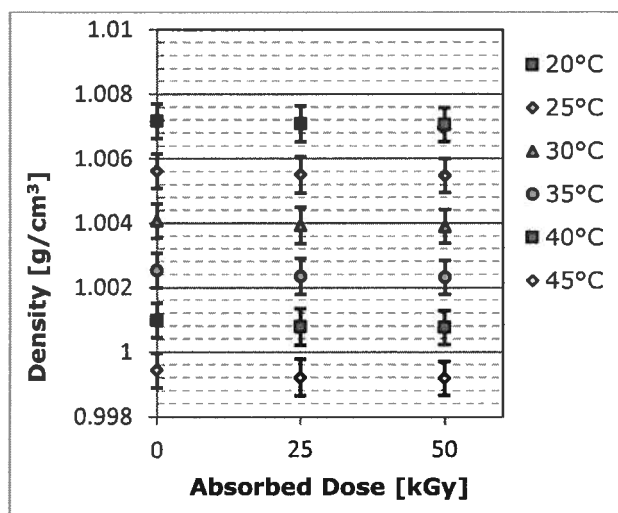


Fig 2 Density of 0.25M HNO<sub>3</sub> as a function of absorbed dose at low dose rate

### 3.2. VISCOSITY

Viscosity measurements were performed using a *KPG-UBBELOHDE micro-viscometer* (0.32mm diameter capillary) on approximately 2.7mL of fresh and irradiated solutions keeping the samples in a thermostatic bath with temperature resolution of  $\pm 0.1^\circ\text{C}$ . The samples were analysed at temperature ranging from  $20^\circ\text{C}$  to  $45^\circ\text{C}$ , with incremental temperature of  $5^\circ\text{C}$ . In order to perform accurate measurements, an equilibration time of about 10 minutes was applied at each temperature step to enable the sample to reach a constant temperature. A digital stopwatch with time resolution of 1/100 second was used to measure the time interval the leading edge of the meniscus of the sample takes to descent from the upper timing mark to the lower one. The Hagenback correction was subtracted from the measured efflux time to obtain the corrected efflux time. Finally, the kinematic viscosity was obtained by multiplying this value for a constant coefficient K, typical of the micro-viscometer used.

The kinematic viscosity is the mean of 5 measurements, with an accuracy of  $\pm 0.7\%$ , including the Hagenback correction and the errors related to temperature variations and micro-viscometer set-up. The estimation of the total measuring uncertainty has been carried out according to Minimum Mean Square Error method [16].

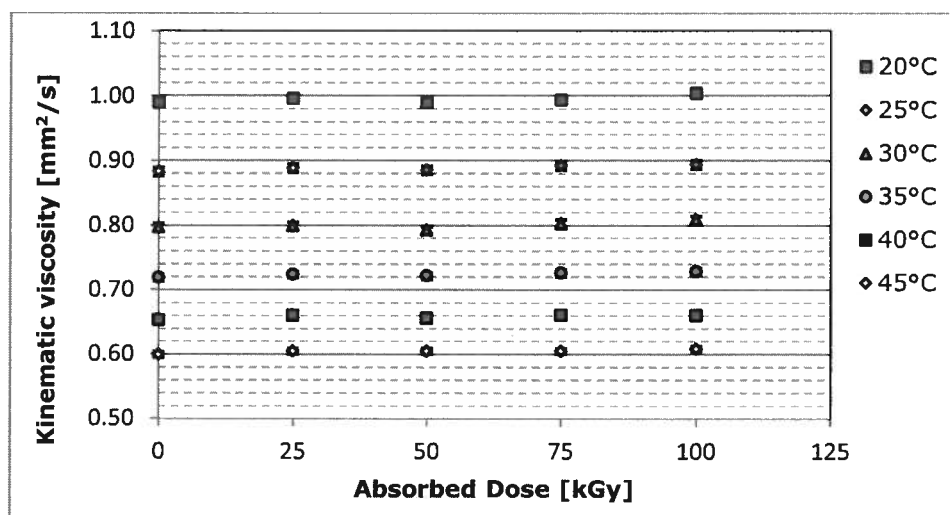


Fig 3 Kinematic viscosity of 0.25M HNO<sub>3</sub> as a function of absorbed dose at high dose rate

The results obtained for irradiated 0.25M HNO<sub>3</sub> samples are reported in Fig 3 and in Fig 4. No kinematic viscosity radiation-induced variation is noticeable for the absorbed dose range considered within the limit of experimental error at both dose rates. Similar trends were found for irradiated 0.5M HNO<sub>3</sub> solutions.

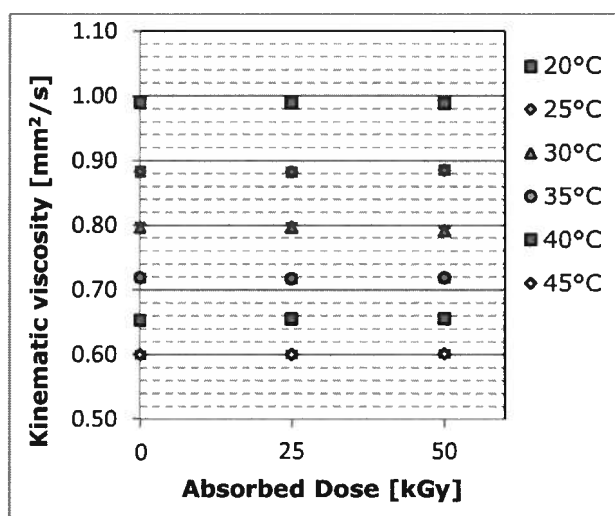


Fig 4 Kinematic viscosity of 0.25M HNO<sub>3</sub> as a function of absorbed dose at low dose rate

### 3.3. ACIDITY

Total acidity of approximately 5mL of 0.5M HNO<sub>3</sub> and 10mL of 0.25M HNO<sub>3</sub> was measured before and after irradiation by titration with 0.1M NaOH. Methyl orange was used as indicator to determine the end point of the titration. Acidity was measured with an uncertainty of  $\pm 0.45\%$ .

Table 1 H<sup>+</sup> concentration of 0.25M and 0.5M HNO<sub>3</sub> as a function of absorbed dose at high dose rate

	0 kGy	25 kGy	50 kGy	75 kGy	100 kGy
<b>0.25 M</b>	0.250 $\pm$ 0.001	0.250 $\pm$ 0.001	0.250 $\pm$ 0.001	0.250 $\pm$ 0.001	0.251 $\pm$ 0.001
<b>0.5 M</b>	0.506 $\pm$ 0.002	0.507 $\pm$ 0.002	0.505 $\pm$ 0.002	0.504 $\pm$ 0.002	0.505 $\pm$ 0.002

Table 2 H<sup>+</sup> concentration [M] of 0.25M and 0.5M HNO<sub>3</sub> as a function of absorbed dose at low dose rate

	0 kGy	25 kGy	50 kGy
<b>0.25 M</b>	0.250 $\pm$ 0.001	0.251 $\pm$ 0.001	0.251 $\pm$ 0.001
<b>0.5 M</b>	0.499 $\pm$ 0.002	0.498 $\pm$ 0.003	0.499 $\pm$ 0.003

The results obtained for 0.25M and 0.5M HNO<sub>3</sub> samples at both high and low dose rate are reported in Table 1 and in Table 2. No radiation-induced modifications are noticeable within the considered irradiation conditions.

### 3.4. NITRATE ANION CONCENTRATION

It is known from the literature that in diluted nitric acid solution the reaction between nitrate ions and water radiolytic products leads to nitrite ion formation. Moreover, an increase of nitrate ion concentration induces increasing nitrite ion production [18]. Nitrite ion is the most important radiolytic product of irradiated nitric acid solution because it directly affects the redox chemistry of several actinides [17]. For this reason, in this work the nitrate ion concentration was monitored by two different techniques.

### 3.4.1. UV spectrophotometry

UV/VIS spectrophotometric analyses of fresh and irradiated  $\text{HNO}_3$  solutions were performed by Lambda EZ210 UV/VIS spectrophotometer (PerkinElmer) in the wavelength range of 230–600nm. All solutions were diluted with the same ratio (1:2) in order to read an optimal absorbance response and to introduce the same error due to dilution. Nitrate anion peak was recognised at 301nm. A calibration of nitrate anion concentration was obtained by measuring the  $\text{NO}_3^-$  absorbance in solutions of  $\text{HNO}_3$  ranging from 0.15M to 0.6M, with increasing concentration of 0.05M.

The nitrate anion concentrations measured for 0.25M and 0.5M  $\text{HNO}_3$  samples are reported in Table 3 and in Table 4. In the case of 0.25M  $\text{HNO}_3$  no modifications were observed, while, in the case of 0.5M  $\text{HNO}_3$  a nitrate anion consumption of about 4.5% at 100kGy with high dose rate and of about 4% at 50kGy with low dose rate were observed.

**Table 3 Nitrate anion concentration in 0.25M and 0.5M  $\text{HNO}_3$  as a function of absorbed dose at high dose rate**

	0 kGy	25 kGy	50 kGy	75 kGy	100 kGy
<b>0.25 M</b>	0.242±0.001	0.243±0.001	0.244±0.001	0.244±0.001	0.241±0.001
<b>0.5 M</b>	0.526±0.005	0.516±0.005	0.513±0.005	0.502±0.005	0.503±0.005

**Table 4 Nitrate anion concentration in 0.25M and 0.5M  $\text{HNO}_3$  as a function of absorbed dose at low dose rate**

	0 kGy	25 kGy	50 kGy
<b>0.25 M</b>	0.250±0.001	0.250±0.001	0.249±0.001
<b>0.5 M</b>	0.523±0.005	0.505±0.005	0.502±0.005

### 3.4.2. RAMAN spectroscopy

In order to further prove the results of UV/VIS spectrophotometric analyses, Raman spectra of approximately 0.2mL of fresh and irradiated  $\text{HNO}_3$  in corked glass NMR tubes were acquired at room temperature. Each spectrum was collected as averaged result of four acquisitions in the range of 100–3800 $\text{cm}^{-1}$ . Spectra analyses were carried out using *Omnic 7.1* software.

The experimental apparatus consists of Ar<sup>+</sup> Laser Stabilite2017 from Spectra-Physics, 514.5nm excitation wavelength, 75mWatt power on sample; Labram HR800 Raman spectrometer from Horiba Jobin Yvon, coupled with Olympus BX41 microscope, 20X lens focal, 1024 pixel CCD thermo-electrically cooled, notch filter, 1800 groove/mm grating.

In all the spectra analysed water stretching, water bending, nitrate and nitrite anion peaks are easily recognised [19]. In order to perform a semi-quantitative evaluation of the nitrate ion concentration, water bending peak was assumed as internal reference in both fresh and irradiated samples and the ratio between the nitrate ion peak and the water bending peak net areas were calculated. As shown in Fig 5 and in Fig 6, the ratios obtained for 0.25M and 0.5M  $\text{HNO}_3$  samples remain constant within the considered irradiation conditions. Therefore, no radiation-induced modification seem to occur. The Raman results seem in contradiction with the evidences of the UV/Vis spectrophotometric measurements concerning the 0.5M  $\text{HNO}_3$  series. In order to achieve a deeper understanding of this phenomenon, further investigations are already in progress, even resorting to ATR FT-IR spectroscopy and Ion Chromatography [20].

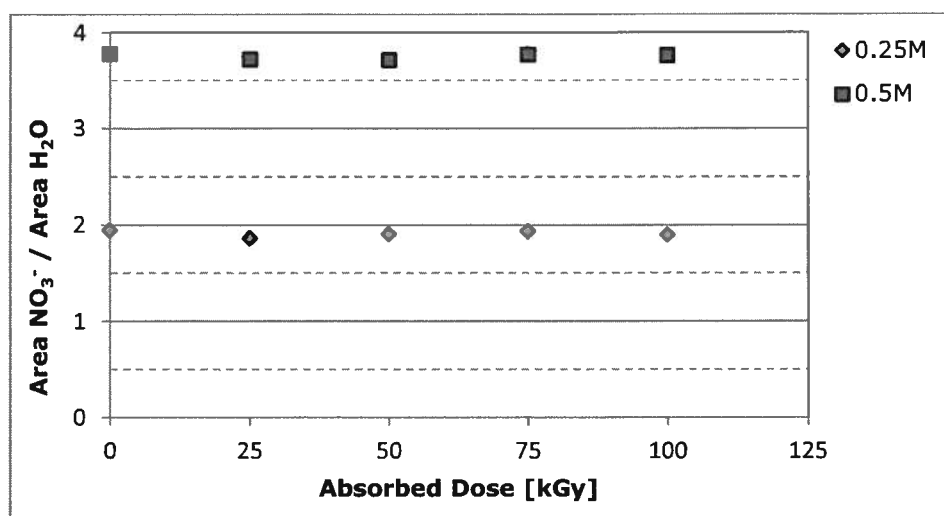


Fig 5 Ratio between nitrate anion and water bending peaks net areas of 0.25M and 0.5M HNO<sub>3</sub> as a function of absorbed dose at high dose rate

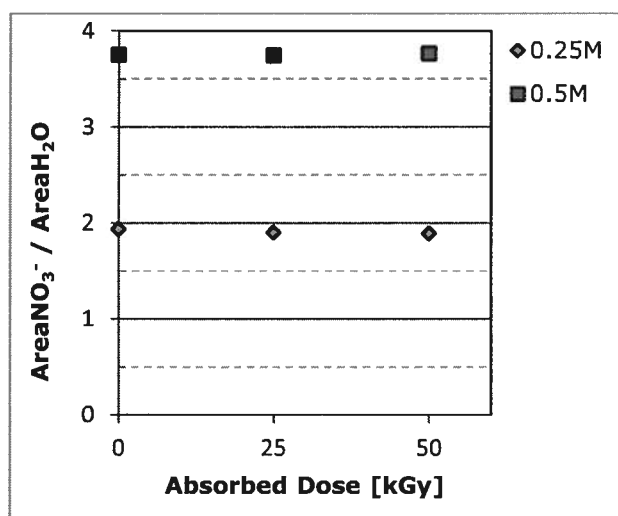


Fig 6 Ratio between nitrate anion and water bending peaks net areas of 0.25M and 0.5M HNO<sub>3</sub> as a function of absorbed dose at low dose rate

## 4. CONCLUSION

In this preliminary study no variation of macroscopic physico-chemical properties was observed in irradiated solutions with increasing the absorbed dose in the range considered with respect to the un-irradiated reference sample. Therefore, no significant changes in the fluid-dynamics of the extracting system are expected. This result is very encouraging for a future industrial-scale development. Concerning the nitrate anion concentration, a few percent consumption in the most concentrated solution is observed only by UV spectrophotometric technique and further investigations involving other spectroscopic and chromatographic techniques are already in progress.

The study of radiation-induced modification of physico-chemical properties will be extended from the aqueous diluent without the extracting ligand to the complete aqueous solvent and, finally, to the whole system in which the hydrophilic ligand containing aqueous phase will be irradiated in contact with the organic phase. The analyses on the irradiated complete aqueous solvent proposed for the selective stripping of actinides are already in progress.

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